

Understanding the Palladium Hydrogen (Deuterium) Electrochemistry as Crucial Step to Approach Low Energy Nuclear Reactions

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Properties of Nanostructures

Rapid insertion reactions (fast insertion and removal of H into and from bulk -Pd)

- adsorption- and absorption process distinguishable in the voltammetry





Octaethylene glycol monododecyl ether ($C_{12}EO_8$)



Octaethylene glycol monohexadecyl ether ($C_{16}EO_8$)

Phase diagram



Templated Electrodeposition:



The porous structure of the metal film



Nanostructured platinum films:

formed by electrochemical deposition from lyotropic phases

Hexagonal H_I

Cubic V_I

SEM Image





50 nm



– 50 nm

Science, 278, 838 (1997).

SEM of nanostructured Pd



Deposited from H₁ lyotropic liquid crystalline phase $(12\% (NH_4)_2 PdCl_4, 47\% C_{16}EO_8, 39\% water, 2\%$ heptane by weight) at 0.1 V vs SCE.

TEM and XRD of nanostructured palladium



Deposited from H₁ lyotropic liquid crystalline phase $(12\% (NH_4)_2 PdCl_4, 47\% C_{16}EO_8, 39\%$ water, 2% heptane by weight) at 0.1 V vs SCE.

TEM of nanostructured palladium



Deposited from H₁ lyotropic liquid crystalline phase $(12\% (NH_4)_2 PdCl_4, 47\% C_{16}EO_8, 39\%$ water, 2% heptane by weight) at 0.1 V vs SCE.

TEM of nanostructured rhodium



Deposited from H₁ lyotropic liquid crystalline phase (12 wt% RhCl₃, 47 wt% $C_{16}EO_8$, 39 wt% water, 2 wt% heptane) at -0.2 V vs SCE

Electrochemical properties of nanostructured palladium



 $\begin{array}{l} 20 \text{ mV/s} \\ 1\text{M} \text{ H}_2\text{SO}_4 \end{array}$

a) 0.002 cm² electrode area

b) 0.54 cm² surface area, corresponds to 27.6 m²/g and 10^7 cm²/cm³

Hydrogen electrode reaction

$$\begin{array}{l} H_{3}O^{+}+e^{-} \rightarrow H_{ads}+H_{2}O \ \ Volmer \\ \\ H_{ads}+H_{3}O^{+}+e^{-} \rightarrow H_{2}+H_{2}O \ \ Heyrovsky \\ \\ H_{ads}+H_{ads} \ \ \rightarrow \ \ H_{2} \ \ \ Tafel \end{array}$$

Additives affecting the H - adsorption



H - insertion into Pd



Pd - Hydride phases



H was loaded into bulk-Pd at different potentials. The electrode was held at each potential for 120s and swept back to -0.2V.

Phys. Chem. Chem. Phys., 2002, 4, 3835



Increasing crystal violet coverage on Pd surface

 H_1 -e Pd initially deposited on gold (0.0079 cm²)

Increasing Pt coverage on Pd surface



10 mV/s 1 M H₂SO₄

3.5 mC deposition charge

 H_1 -e Pd initially deposited on gold (0.0079 cm²)

Additives affecting the D - adsorption



20 mV/s $D_2O + 0.1 \text{ M Li}_2SO_4$

Electrolysis



Conclusions

- H-absorption occurs without passing through adsorbed state
- Hydrogen diffuses directly into Pd bulk

$$H_3O^+ + e^- \rightarrow H_{abs} + H_2O$$

• Process speeds up when formation of adsorbed hydrogen layer is suppressed by the coverage of poisons

Our presence

- environmental viewpoint millions of tons toxic gases released on every day global warming
- economical viewpoint strongly dependent on fossile fuels oil, coal and gas price rises up

Our options

- solar energy
- wind power
- batteries and fuel cells
- nuclear power

What about cold fusion?

• Pons - Fleischmann experiment in 1989

- long term electrolysis in a heavy water solution on plane palladium as electrode

- Research over the last 18 years
 - excess heat, nuclear transmutation, change of electrode surface morphology during long term electrolysis
 - not only restricted to Palladium but also occurs when nickel, gold etc. used as electrode
- *Journal of New Energy, Fusion Information Centre, Salt Lake City*

Lack of reproducibility!!!



Dr. Marwan Chemie Forschung & Entwicklung

Electrochemical Cell





hanging the PERSPECTIVE

